

Properties of polypyrrole polyvinilsulfate films for dual actuator sensing systems

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ABSTRACT

One of the challenges of modern science is the development of actuators able to sense working conditions while actuation, mimicking the way in which biological organs work. Actuation of those organs includes nervous (electric) pulses dense reactive gels, chemical reactions exchange of ions and solvent. For that purpose, conducting polymers are being widely studied. In this work the properties of self-supported films of the polypyrrole:polyvinilsulfate (PPy/PVS) blend polymer were assessed. X-ray photoelectron spectroscopy (XPS) studies show how during reduction / oxidation the polymer exchanges cations when immersed in a NaClO₄ aqueous solution, revealing free positive charges in the electrolytic solution as the driving agents leading to the swelling/shrinking of the polymer. Eventually it is the phenomenon responsible of the actuation of the polymeric motors. Submitting the system to consecutive potential sweeps shows the reaction is really sensing the scan rate used in each cycle revealing that while actuating the system is actually sensing the electrochemical working conditions.

Keywords: conducting polymers, electro-chemo-biomimesis, artificial muscles, polypyrrole

1. INTRODUCTION

An uninterrupted feedback communication about the working conditions between some organs and the brain is what originates proprioception^{1,2}. Living cells possess reactant macromolecules that through reactions (molecular motors) develop conformational movements³.

The last decades have witnessed the development of a plethora of electrochemical devices such as conducting polymers or graphenes⁴⁻⁶. When these reactions take place the relative proportions of ions, solvent and polymer changes leading to conformational movements, as it happens inside the biological cells^{7,8}. These composition dependent properties are allowing the creation of new devices such as artificial muscles^{9,10}, smart windows^{11,12} or biosensors¹³. Sensing the working conditions during actuation these artificial systems mimic biological ones¹⁴. In a physically uniform device several sensors and an actuator work simultaneously. The source of this dual behavior is the electrochemical reaction.

This work studies the electrogeneration and electrochemical properties of polypyrrole/polyvinilsulfate (PPy/PVS) films after having been removed from the metal electrode over which they were generated, as well as their sensing properties when submitted to potential sweep rates the scan rate varies.

2. METHODOLOGY AND RESULTS

2.1 Chemicals and Equipment

To electrogenerate the polymeric blend, the following substances were employed: distilled pyrrole (Fluka), Poly(vinylsulfonic acid, sodium salt) solution (NaPVS) (Sigma-Aldrich, 25 wt% in H₂O), lithium and sodium perchlorate (Sigma-Aldrich) and Milli-Q water.

An Autolab potentiostat/galvanostat PGSTAT302 (NOVA 1.11 software) was employed. A one-compartment electrochemical cell with an Ag/AgCl (3M KCl) reference electrode (Metrohm), two stainless steel AISI 316 with a total immersed area of 14.56 cm² as counter electrodes surrounding a working electrode of the same material with a total immerse surface of 7.28 cm²

A Sartorius Extended ED224s balance was used to weigh the films and an electronic microscope Hitachi S.3500N was employed for X-ray photoelectron spectroscopy (XPS).

2.2 Film Preparation

PPy/PVS films were generated at room temperature over the working electrode from a 0.05M NaPVS and 0.1M pyrrole aqueous solution. But first a study was carried out to find suitable electrochemical conditions to synthesize the material.

The potential window of the electrolyte (0.05M NaPVS aqueous solution) was found by subjecting the solution to potential sweeps at 40 mVs^{-1} in the potential interval (-1200 mV, 1000mV). Figure 1 shows how if the interval is wider than (-1100 mV, 900 mV) the water electrolysis begins, either with hydrogen release if the cathodic limit is open or with oxygen release on the anodic side.

If the cycling is carried out with the 0.05M NaPVS and 0.1M pyrrole aqueous solution the voltammetric responses reveal that the monomeric oxidation polymerization takes place for potentials over 650 mV with a clean working electrode and over 400 mV with a partially coated with monomer electrode. In Figure 1 it can also be seen that the first cycle exhibits a loop, indicating that the process start by nucleation. The oxidation-polymerization is produced between 400 mV and 900 mV, far enough from the limits where the water electrolysis started.

With this information, to electrogenerate the polymer blend, a potential of 800 mV was first applied until a uniform film coated the metal and a charge of 0.5C was consumed. After that, the potential was lower to 500 mV and kept for 37000s. This time was considered enough for the films to be removed from the metal. The mass of the rinsed and dry film generated this way was of 12.0 mg

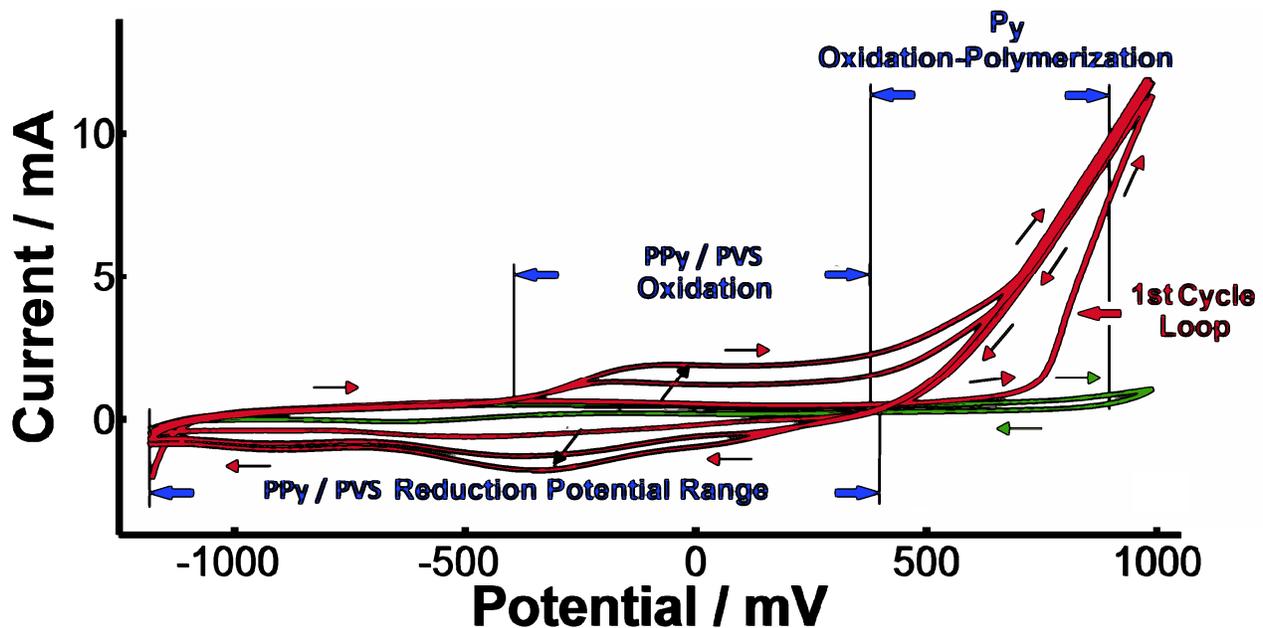


Figure 1. Voltammetric responses of the background electrolyte (green) (0.05 M NaPVS aqueous solution) and the three first consecutive cycles in a 0.05M NaPVS and 0.1M pyrrole aqueous solution (red), at a scan rate of 40 mVs^{-1} .

2.3 Influence of the Cathodic Limit

The Figure 2 shows the results of subjecting the film to a potential sweeps rates between a fixed anodic potential of 700 mV and a cathodic limit varying from - 600 mV to -1800 mV in a 0.1 M LiClO_4 aqueous solution. Stationary voltammetric responses (reached after two cycles) display a shift of the anodic peak towards more positive potentials.

This phenomenon can be explained by the appearance of structures like lamellas, vesicles and micelles¹⁵, more stable from an energetic point of view. Since the only source of energy available is the potential (the experiments were realized at constant temperature, electrolyte concentration and so on,) to reoxidize these deep reduced formations a higher potential is needed¹⁶⁻¹⁸. Due to a complete oxidation of the film erases its memory the reduction peak appears always at the same potential.

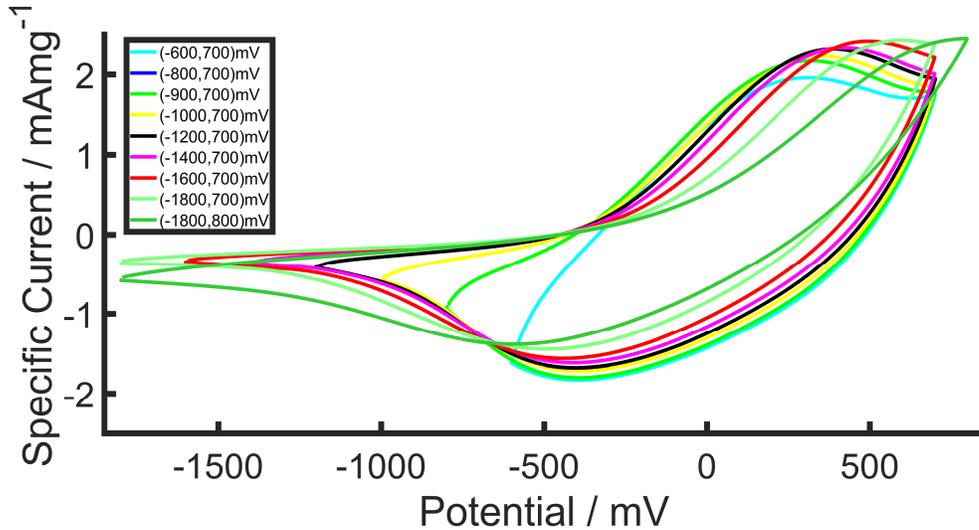


Figure 2. Voltammetric responses corresponding to a self-supported PPy/PVS immersed in a 0.1M LiClO₄ aqueous solution with a fixed anodic limit of the 700 mV and a cathodic one going from -600 mV to -1800 mV. Realized with a scan rate of 10 mVs⁻¹.

2.4 Ionic Exchange

To determine which kind of ions this polymeric blend exchanges when working immersed in a 0.1 M NaClO₄ aqueous solution, three films were analyzed by XPS, in three different states, right after synthesis, after voltammetric reduction and after voltammetric oxidation. The results (Figure 3) shows a higher cation (Na⁺) content in the reduced sample indicating that the polymer is changing cations following the reaction



Being PPy the active centers in the polymer that will accommodate a positive charge after oxidation, PVS⁻ the anion that was trapped during the synthesis in the film and Na⁺ the exchanged cation during the redox process.

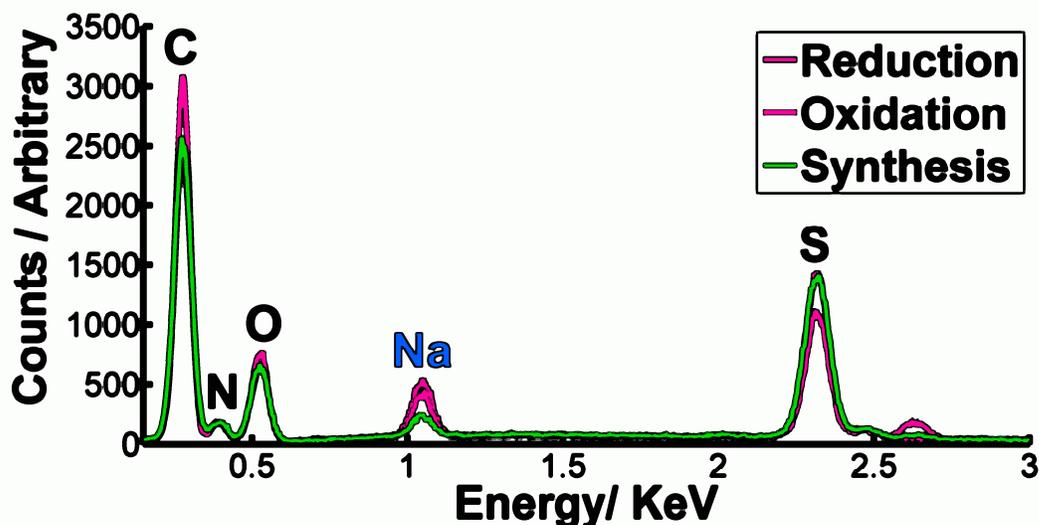


Figure 3. XPS analysis of three PPy/ PVS films in three different states, right after synthesis, oxidized and reduced.

2.5 Influence of the Scan Rate

After the study of some electrochemical properties of the PPy/PVS films, their sensing properties were assessed by changing the scan rate during cycling voltammetry experiences. Figure 4a shows the stationary voltammetric (after two cycles) responses attained for scan rates ranging from 8 mVs^{-1} to 100 mVs^{-1} between the potential limits of -800 mV and 500 mV . In the graphic it can be seen clearly a scan rate (20 mVs^{-1}) above which the reduction and oxidation peaks have disappeared. Focusing on the lowest scan rates where these maxima are clearly depicted it can be appreciated how they separate from each other when the scan rate increases, shifting the oxidation peak to more anodic potentials and the cathodic peak to more cathodic. This phenomenon indicates the existence of a resistance to the entrance and exit of cations that raises with the scan rate.

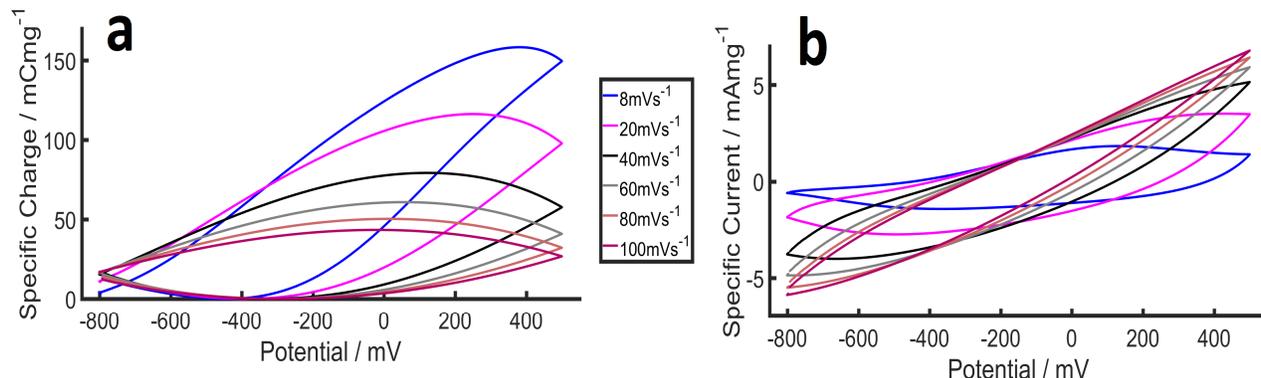


Figure 4. (a) Voltammetric responses for a self-supported PPy/PVS at different scan rates, from 8 mVs^{-1} to 100 mVs^{-1} in the potential window $(-800, 500) \text{ mV}$ (b) Concomitant coulouvoltammetric responses.

Through integration of the voltammetric responses of Figure 4a, the coulouvoltammetric ($Q-E$) are obtained. The close loop shown in the most part of these responses represent a region where the reduction positive charge equals the oxidize negative charge. The difference between the maximum and the minimum point of this region represents the reversible charge implicated in the oxidation/reduction of the polymer blend. The left open part represent the irreversible charge associated to other reactions like hydrogen evolution¹⁵.

Figure 5 shows that, in fact, both charges decrease with the scan rate. Since other parameters such as temperature or electrolyte concentration were kept constant during the experiment this reduction in the charge values with the scan rates indicate the presence of less electrochemical energy available per potential sweep for all reactions to take place.

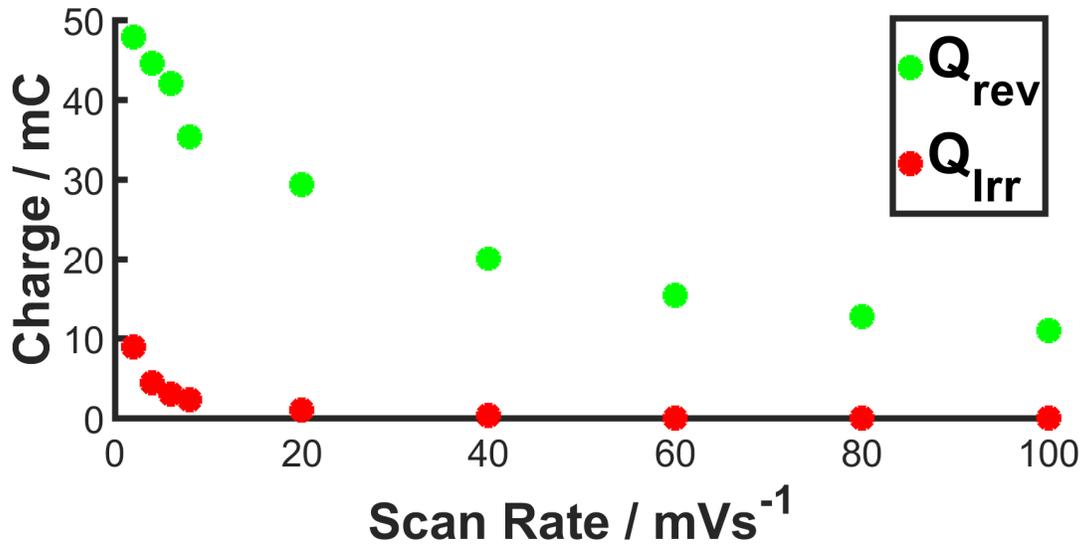


Figure 5. Evolution of the charge (reversible and irreversible with the scan rate)

As mentioned, a part from the scan rate the other parameters (pressure, temperature,...) remain constant. In these conditions the effective electrochemical energy (Δu) at which the film is subject can be expressed as

$$\Delta u = \frac{\Delta U(J)}{v(Vs^{-1})} = \Delta Q(C) \frac{\Delta E(V)}{v(Vs^{-1})} = \frac{\Delta Q \cdot \Delta E}{v} (Cs) \quad (2)$$

Where v is the scan range, E the potential window and Q , the reversible consumed charge, that according to the electrochemically stimulated conformational relaxation (ESCR) model is related to v through an exponential relation. The experimental results follow the equation

$$\Delta u = a e^{bv} \quad (3)$$

Where $a=11.14 \text{ mJ mV}^{-1} \text{ s}$, $b=-0.084 \text{ mC}^{-1}$, and the correlation coefficient is $r^2=0.99$ (Figure 6)

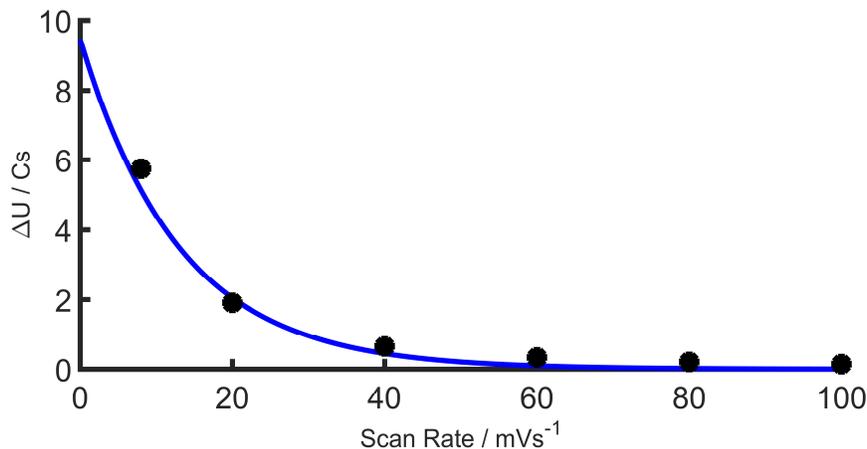


Figure 6. Evolution of the effective consumed electrical energy

The ESCR model establishes that there should be two exponential functions, one related to the conformational movements of the polymeric chains and other with the ion diffusion. What it happens in this case is that the scan rate range chosen is too high for the conformational oxidation/compaction and reduction/relaxation have enough time to occur. Under the conditions used the reaction controlled by diffusion.

The important consequence of this is that the electrical energy is actually sensing the working conditions, the scan rate in this case

3. CONCLUSIONS

Polymeric films of PPy/PVS thick enough to be peeled from the electrode were successfully generated and electrochemically studied showing indications of structures like lamelles or vesicles when the cathodic limit increases and the film is more and more reduced. This originates an anodic shift of the oxidation peak due to the need of reaching higher potentials to get the necessary energy to reoxidize. XPS analysis shows a cation exchange with the electrolyte. It was also proved that the electrochemical reaction is able to sense, through the electrical energy, the working conditions

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